

KENYAPLEX EXAMS 2019

FORM 3

CHEMISTRY PAPER 3 PRACTICAL

CONFIDENTIAL.

In addition to the apparatus and fittings found in the laboratory, each student will require the following:

1. About 80cm³ of solution A
2. About 100cm³ of solution B
3. About 70cm³ of solution C
4. 1 pipette
5. 1 burette
6. 3 conical flasks (250ml)
7. A 250ml volumetric flask
8. 1 thermometer (-10° C to 110°C)
9. 8 test tubes
10. 2 boiling tubes
11. 10ml measuring cylinder
12. 7 labels
13. a test-tube holder
14. Solid G (about 0.3g)
15. Solid T (about 0.3g)
16. Glass rod
17. Metallic spatula
18. Solid sodium hydrogen carbonate (about 0.2g)
19. 500 ml distilled water

Access To:

1. Bunsen burner
2. methyl orange indicator supplied with a dropper
3. Bromine water supplied with a dropper
4. 2M sodium hydroxide supplied with a dropper
5. Aqueous Barium nitrate supplied with a dropper
6. 2M Nitric (v) acid supplied with a dropper
7. universal indicator supplied with a dropper
8. PH scale chart.
9. Acidified potassium manganate (vii) supplied with a dropper.
10. Acidified potassium dichromate (vi) supplied with a dropper.

NB:

- Solution A is prepared by dissolving 55ml of concentrated sulphuric (vi) acid in one litre of solution.
- Solution B is prepared by dissolving 8g of anhydrous sodium carbonate in one litre of solution Sodium C is prepared by dissolving 80g of sodium hydroxide in one litre of solution.
- Bromine water is prepared by dissolving 1cm³ of 20 volumes bromine water in 100cm³ of solution.

- Acidified potassium manganate (vii) is prepared by dissolving 3.16g of KMnO_4 in 600cm^3 of $2\text{MH}_2\text{SO}_4$ and made to one litre solution.
- Acidified potassium Dichromate (vi) is prepared by dissolving 6g of $\text{K}_2\text{Cr}_2\text{O}_7$ in 600cm^3 of $2\text{MH}_2\text{SO}_4$ and made to one litre solution.
- 2M bench reagent of Sodium hydroxide is prepared by dissolving 80g of sodium hydroxide in one litre of solution.
- Nitric (v) acid (2) is prepared by dissolving 126ml in one litre of solution.
- Barium nitrate solution is prepared by dissolving 0.05g in one litre of solution.
- Solid G = hydrated sodium carbonate
- Solid T = Maleic acid.