

KENYAPLEX EXAMS
Term 1
233/3 CHEMISTRY PRACTICAL
CONFIDENTIAL INSTRUCTIONS

Instructions to Schools:

The information contained in this paper is to enable the head of the school and the teacher in charge of Chemistry to make adequate preparation for the Chemistry Practical Examination.

NO ONE ELSE should have access to this paper or acquire knowledge of its content. Great care MUST be taken to ensure that the information herein does NOT reach the candidates either directly or indirectly. The teacher in charge of Chemistry should NOT perform any of the experiments in the SAME room as the candidates nor make the results of the experiment available to the candidates or give any information related to the experiments to the candidates. Doing so will constitute an examination irregularity.

REQUIREMENTS FOR CANDIDATES

In addition to fittings, and apparatus found in the chemistry laboratory, each candidate will require:

1. **Solid P** - 4.0 g of accurately weighed ethane-1,2-dioic acid in to stoppered container.
2. 100 cm³ of **Solution Q** – 0.2M sodium hydroxide solution.
3. About 2g of **Solid T** – a mixture of 1.5 g lead (II) carbonate and 0.5 g copper (II) sulphate.
4. **Liquid U** - 5 ml of ethanol in a stoppered boiling tube.
5. About 0.5g of sodium hydrogen carbonate in a stoppered container.
6. Distilled water in a wash bottle.

7. One 50 cm³ burette.
8. One 25 cm³ pipette.
9. One 10 cm³ measuring cylinder.
10. One empty boiling tube.
11. Six, clean dry test tubes placed in a rack.
12. One thermometer with -10°C to 110°C range.
13. One test tube holder.
14. One clean and dry conical flask.
15. A white tile.
16. A watchglass.

ACCESS TO:

1. Phenolphthalein indicator supplied with a dropper.
2. 2M HCl supplied with a dropper.
3. 2M nitric (V) acid supplied with a dropper
4. Aqueous NaOH supplied with a dropper.
5. 2M Ammonia solution supplied with a dropper
6. Source of heat (Bunsen burner).
7. Acidified potassium dichromate (VI) – K₂Cr₂O₇.

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