

KENYAPLEX EXAMINATION -2019

END OF TERM 1 EXAM

CHEMISTRY PAPER 3

FORM 4

CONFIDENTIAL

In addition to the apparatus found in the laboratory each candidate will require the following;

- About 1g of solid G
- 6 clean test-tubes
- Universal indicator solution and a pH chart
- Ethanol supplied with a dropper
- Clean dry metallic spatula
- 1 boiling tube
- Distilled water
- Solution J, about 130cm³
- Solution Q, about 160cm³
- Solution R, about 30cm³
- Screened methyl orange indicator
- Methyl orange indicator
- 100ml measuring cylinder
- Filter paper
- Means of labeling
- Solid P
- Thermometer
- 100ml beaker

Access to the following;

- ❖ Ethanol supplied with a dropper
- ❖ Concentrated sulphuric (VI) acid supplied with a dropper bottle
- ❖ Acidified Potassium dichromate (VI) solution
- ❖ Acidified Potassium Manganate (VII) solution.
- ❖ 2M Ba(NO₃)₂ solution.
- ❖ 2M NaOH solution.
- ❖ 2M HCl acid.
- ❖ Source of heat.

Preparation

Solution J is 0.12M HCL, prepared by adding about 800cm³ of distilled water to 4.05cm³ of concentrated HCL of density 108g/cm³ and making it to one litre of solution

Solution Q is prepared by dissolving 5.3g of anhydrous sodium carbonate in enough distilled water and making up to one litre of solution.

Solution R is prepared by dissolving 15.75g of hydrated barium hydroxide in enough distilled water and top up to one litre of solution.

Solid P is 2.0g of oxalic acid weighed accurately and supplied in a stoppered container

Solid G is sodium sulphite