FORM FOUR CHEMISTRY PRACTICAL.

FEBRUARY 2017

NAME: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ CLASS: \_\_\_\_\_\_\_\_\_\_\_\_\_\_

You are provided with:

* 0.2 M hydrochrolic acid, solution P.
* 5.3 g of a metal carbonate, M2CO3 dissolved in distilled water to form 600 cm3 of solution as solution Q.

You are required to calculate the relative atomic mass of M

Procedure

Fill the burette with hydrochloric acid, solution P. Pipette 25cm3 of solution Q into a 250 ml conical flask.

Add 2-3 drops of phenophalein indicator and titrate with solution P.

Record the results in the table below. Repeat the procedure two more times to complete the table (4mks)

|  |  |  |  |
| --- | --- | --- | --- |
|  | I | II | III |
| Final burette reading |  |  |  |
| Initial burette reading |  |  |  |
| Volume of solution P used(cm3) |  |  |  |
|  |  |  |  |

1. Calculate the average volume of solution P used.(1 mk)
2. Write the equation for the reaction that takes place. (1 mk)
3. Calculate the:
4. Number of moles of solution P in the average titre (1 mk)
5. Number of moles of M2CO3 in 25 cm3 of solution Q. (1 mk)
6. Concentration of Q in moles per litre (Morality). (1 mk)
7. Concentration of Q n grams per litre (1 mk)
8. The relative Formula Mass(RFM) of the M2CO3. (1 mk)
9. Relative Atomic Mass of Metal M. (C=12.0, O=16.0). (1 mk)