**NAME …………………………..……………….. DATE …………………………**

**INDEX NO. ……….……….…………………...…..… SIGNATURE ……………..…………..**

**233/2**

**CHEMISTRY**

**PAPER 2**

**(THEORY)**

**JULY/AUGUST, 2014**

**TIME: 2 HOURS.**

**MBOONI EAST SUB - COUNTY JOINT EVALUATION TEST**

*Kenya Certificate of Secondary Education.*

**233/2**

**CHEMISTRY**

**PAPER 2**

**(THEORY)**

**TIME: 2** **HOURS.**

**INSTRUCTIONS TO CANDIDATES.**

* Write your name and index number in the spaces provided above.
* Sign and write the date of exam in the spaces provided above.
* Answer **ALL** the questions in the spaces provided.
* Mathematical tables and silent electronic calculators may be used.
* All working **MUST** be clearly shown where necessary.
* This paper consists of 12 printed pages. Candidates should check to ensure that all pages are printed as indicated and no questions are missing

**FOR EXAMINER’S USE ONLY.**

|  |  |  |
| --- | --- | --- |
| **Questions** | **Maximum score** | **Candidates score** |
| 1 | 13 |  |
| 2 | 12 |  |
| 3 | 13 |  |
| 4 | 10 |  |
| 5 | 10 |  |
| 6 | 12 |  |
| 7 | 10 |  |
| **Total score** | **80** |  |

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**233/2**

**Chemistry**

**Paper 2 (Theory)**

1. The grid below shows a section of the periodic table. The letters do not represent the actual symbols for elements. Study it and answer the questions that follow.

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| X |  |  |  |  |  |  |  |  |
|  |  |  |  | Y |  | C | W |  |
|  | Z |  | R |  |  |  | A |  |
| E |  |  |  |  |  | D |  | B |
|  |  |  |  |  |  |  |  |  |

(a) Compare the atomic radii of elements:

(i) A and R. Explain (2 marks)

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(ii) C and D. Explain (2 marks)

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(b) Give the name of the family in which elements W and A belong ( ½ mk)

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(c) Identify the element which is

(i) Most electronegative ( ½ mark)

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(ii) Strongest reducing agent. ( ½ mark)

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(d) Write the formula of the oxide of X ( ½ Mark)

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(e) (i) What type of bond exists in the chloride of R (1 Mark)

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(ii) Draw the structure of the molecule of chloride of R in gaseous state (1 Mark)

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(iii) Write an equation for the reaction between element A and water. (1 Mark)

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(f) The table below shows the melting points of some period 3 oxides.

|  |  |  |  |
| --- | --- | --- | --- |
| Oxide | SiO2 | P4O6 | SO2 |
| M.P 0C | 1610 | 23.6 | -72.7 |

Explain the difference in melting points of SiO2 and P4O6. (1 Mark)

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(g) Excess hydrogen gas is reacted with 4.5g of heated oxide of metal Q. 3.6g of Q is finally left in the

reaction vessel. Given that O = 16, Q has a valency of 2.

(i) Write the equation for the reaction taking. (1 Mark)

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(ii) Calculate the relative atomic mass of Q. (2 Marks)

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1. (a) Give the names of the following compounds.

(i) CH3CH2CH2CH2CH2OH (1 Mark)

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(ii) CH3CH2COOH (1 Mark)

.……………………………………………………………………………………………………………...

O

(iii) (1 Mark)

H - C-O-CH2-CH3

.……………………………………………………………………………………………………………...

(b) Describe a chemical test that can be used to distinguish between substances (i) and (ii) above (a)

(1 Mark)

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(c) Study the information in the table below and answer the questions that follow.

|  |  |
| --- | --- |
| Number of carbon atoms per molecule | Relative molecular mass of hydrocarbon |
| 2 | 28 |
| 3 | 42 |
| 4 | 56 |

(i) Write the general formular of the hydrocarbons in the table. ( ½ Mark)

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(ii) Predict the relative molecular mass of the hydrocarbon with 5 carbons. (1 Mark)

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(iii) Determine the molecular formular of the hydrocarbon in (ii) above and draw its structural

formular (H = 1, C = 12) ( 2 Marks)

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(d) Study the scheme below and answer the questions that follow.

C2H5COONa

CH≡CH

C2H6

Complete

combustion

Step IV

CH2 = CHCI

Step II

(CH2-CHCl)n

Step III

Step V

(i) Name the reagents used in

Step I …………………………………………….. ( ½ Mark)

Step II …………………………………………….. ( ½ Mark)

Step IV ……………………………………………. ( ½ Mark)

(ii) Write an equation for the complete combustion of CH≡CH ( 1 Mark)

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(iii) Explain one disadvantage of the continued use of items made from the compound formed in Step III. (2 Marks)

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1. Fractional distillation of liquid air usually produces Nitrogen and Oxygen as the major products.

(a) (i) Name one substance used to remove carbon (IV) Oxide from air before it is changed into liquid. (1 Mark)

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(ii) Describe how nitrogen is obtained from liquid air.

(Boiling point of Nitrogen = -1960C, Oxygen = -1830C) (3 Marks)

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(b) Study the flow diagram below and answer the questions that follow:-

Air

Water

Step 2

Step 1

Electrolysis

N2(g)

M

Step 2

Ammonia

Step 4

Rhodium/Platinum

Temperature 9000C

Colourless gas Q

Step 5

Nitrogen (IV) Oxide

Step 6

Nitric Acid

Step 7

Ammonium Nitrate

(i) Name the element M (1 Mark)

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(ii) Why is it necessary to use excess air in step 4 (1 Mark)

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(iii) Write an equation for the reaction in step 7 (1 Mark)

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(iv) In the harber process, the optimum yield of ammonia obtained is when a temperature of 4500C a

pressure of 200 atmospheres and iron catalyst are used.

N2(g) +3H2(g) 2NH3(g) , Δ = -92kJmol-1

(I) How would the yield of ammonia be affected if the temperature is raised to 6000C. Explain

(2 Marks)

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(II) State two uses of ammonia gas (2 Marks)

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(c) State and explain the observations made if a sample of sulphur is heated in concentrated

Nitric(V) Acid (2 Marks)

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1. 1g of magnesium ribbon was reacted with hydrochloric acid at room temperature in order to investigate how the rate of reaction varies with time. The results obtained were recorded as shown below.

|  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Time (seconds) | 0 | 20 | 40 | 60 | 80 | 100 | 120 | 140 | 160 | 180 |
| Volume of gas produced (cm3) | 0 | 10 | 20 | 26 | 32 | 35 | 38 | 39 | 40 | 40 |

1. (i) On the graph provided, plot a graph of volume of gas produced against time taken. Label the

graph K. (3 Marks)





(ii) From the graph determine the rate of production of the gas at 110 seconds. (2 Marks)

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(b) On the same axis sketch the graph you would expect to obtain if:-

(i) The same mass of powdered magnesium was used instead of magnesium ribbon. Label the graph

Y. (1 Mark)

(ii) If the temperature of the solution mixture was reduced from 250C to 150C. Label the graph Z. (1 Mark)

(c) Determine the mass of magnesium ribbon that remained unreated in this experiment (Mg = 24, Molar

gas volume = 24dm3 at r.t.p) (3 Marks)

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1. A student set up the apparatus shown below to prepare and collect dry carbon (IV) oxide gas.

Dilute H2SO4 acid

Barium

Carbonate

Conc. H2SO4(l)

Water

(a) State a correction for three mistakes in the set up above. (3 Marks)

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1. 30cm3 of carbon (II) oxide gas were reacted with 10cm3 of oxygen gas. The resulting mixture of

gases was bubbled through an alkali. Determine the volume of the residual gas. (3 Marks)

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1. The flow chart below is for the manufacture of Sodium Carbonate by the solvay process. Use it to

answer the questions that follow.

K

L

M

Tower

N

Lime

Kiln

Reaction chamber

Reaction chamber

R

Substance

T

P

Reaction chamber

Q

NH4Cl3(aq)

Limestone

NaHCO3(s)

Na2CO3(s)

CO2(g)

H2O(l)

H2O2(l)

CaCl2(aq)

NH3(g)

Water

(i) Name substance T (1 Mark)

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(ii) Explain the use of water in chamber R (1 Mark)

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(iii) Name the two raw materials required in the manufacture of Sodium Carbonate. (1 Mark)

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(iv) Write an equation for the reaction taking place in chamber M. (1 Mark)

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1. Use the standard electrode potentials given below to answer the questions that follow.

EθV

Ag+(aq) + e- Ag (s) +0.80

CU2+ (aq) + 2e- Cu (s) +0.34

Pb2+(aq) + 2e- Pb (s) -0.13

Zn2+(aq) + 2e- Zn(s) -0.76

1. Select two half-cells which when combined give the lowest workable cell voltage. (Lowest e.m.f)

(1 Mark)

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(b) calculate the e.m.f of the cell formed by combining the two half – cells in a) above (1 Mark)

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c) (i) Select the strongest oxidizing agent. (½ Mark)

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(ii) Select the strongest reducing agent. (½ Mark)

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(d) A cell was set up using lead and zinc electrodes as shown below.

V

Zinc rod

Salt bridge

1M

Zn2+ (aq)

Lead rod

1M

Pb2+ (aq)

(i) Write the half cell equation for the half – cell in which oxidation occurs. (1 Mark)

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(ii) Write the overall cell equation. (1 Mark)

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(iii) What is the role of the salt bridge? (1 Mark)

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(e) An iron cup was electroplated using chromium. The chromium electrode and the iron cup was

thoroughly cleaned and weighed before being dipped into the electrolyte.

(i) Why was it necessary to clean the metals before dipping them into the electrolyte. (1 Mark)

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(ii) A current of 0.75A was passed through the solution for one hour and four minutes. The mass of

chromium deposited on the cup was 0.52g. (1F = 96500C, Cr = 52)

(I) Calculate the quantity of electricity passed (1 Mark)

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.……………………………………………………………………………………………………………...

(II) How many moles of chromium were deposited (1 Mark)

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(III) Calculate the quantity of electricity required to deposit one mole of chromium. (1 Mark)

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(IV) Calculate the number of Faradays required to deposit one mole of chromium and hence deduce

the charge on the ion. (2 Marks)

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1. (a) The diagram below shows spots of pure substances A, B and C on a chromatography paper. Spot D is

that of a mixture.

A

B

C

D

After development A, B and C were found to have moved 7cm, 3cm and 5cm respectively. D had separated into two spots which moved 5cm and 7cm.

(i) On the diagram

(I) Label the baseline (1 Marks)

(II) Show the position of the all the spots after development (2 Marks)

(ii) Identify the substances present in mixture D. (1 Mark)

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(b) Describe how solid ammonium chloride can be separated from a solid mixture of ammonium

chloride and sodium chloride. (2 Marks)

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(c) The table below shows liquids that are miscible and those that are immiscible.

|  |  |  |
| --- | --- | --- |
| Liquid | Y | Z |
| W | Miscible | miscible |
| X | Miscible | Immiscible |

Use the above information to answer the questions that follow.

(i) Name the method that can be used to separate W and Y of a mixture of the two ( ½ Mark)

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(ii) Describe how a mixture of X and Z can be separated. (2 Marks)

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(d) Crude oil is a source of many compounds that contain carbon and hydrogen only.

(i) Name the processes used to separate components of crude oil. ( ½ Mark)

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(ii) On what physical property of the above components does the separation depend? ( 1 Mark)

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