NAME \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ INDEX NO. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

SCHOOL \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ SIGNATURE \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ DATE \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**233/3**

**CHEMISTRY**

**PRACTICALS**

**PAPER 3**

JULY/AUGUST, 2017

**TIME: 2¼ HOURS**

233/3

CHEMISTRY

PRACTICALS

PAPER 3

TIME: 2¼ HOURS

**INSTRUCTIONS TO CANDIDATES**

1. Write your name, school and index number in the spaces provided above.
2. Sign and write the date of the examination in the spaces provided above.
3. Read the questions carefully
4. Follow the instruction for each procedure carefully
5. You are not allowed to start working with the apparatus for the first 15 minutes of 2 ¼ hrs allowed for this paper. This time will enable you read the question paper and make sure you have all chemicals and apparatus that you may need.
6. All working must be clearly shown.
7. Mathematical tables and electronic calculators may be used.

FOR EXAMINERS’ USE ONLY:

|  |  |  |
| --- | --- | --- |
| QUESTION | MAXIMUM SCORE | CANDIDATES SCORE |
| 1 | 21 |  |
| 2 | 12 |  |
| 3 | 07 |  |
| TOTAL SCORES | 40 |  |

*This paper consists of 8 printed pages.*

*Candidates should check carefully to ascertain that all the pages are printed as indicated and no questions are missing.*

1. You are provided with:

- Solid V

- 2.0M hydrochloric acid, solution B

- 0.1M sodium hydroxide, solution C

You are required to determine the enthalpy change H, for the reaction between solid V and one mole of hydrochloric acid.

**Procedure I**

Using a burette, place 20.0cm3 of 2.0M hydrochloric acid, solution B in a 100ml beaker. Measure the temperature of the solution after every half – minute and record the values in table 1. At exactly 2½minutes, add all of solid V to the acid. Stir the mixture gently with a thermometer of mixture after every half minute and record the values in table I (Retain the mixture for use in procedure II)

**Table 1**

|  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Time (min) | 0 | 1 | 1½ | 2 | 2½ | 3 | 3½ | 4 | 4½ | 5 |
| Temperature |  |  |  |  |  |  |  |  |  |  |

(4 marks)

a) On the grid provided plot a graph of temperature (vertical axis) against time. (3 marks)



b) From the graph determine the change in temperature T. (1 mark)

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c) Calculate the heat change for the reaction (Assume that the specific heat capacity of the mixture is4.2j/g/K and the density of the mixture is 1g/cm3) (1 mark)

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**Procedure II**

Rinse the burette thoroughly and fill it with 0.1M sodium hydroxide, solution C. Transfer all the contentsof the 100ml beaker used in procedure I into a 250ml volumetric flask. Add distilled water to make up to the mark. Label this solution V. Using a pipette and a pipette filler, pipette 25.0cm3 of solution V into a conical flask. Add 2 – 3 drops of phenolphthalein indicator and titrate with solution C. Record your results in table 2 below. Repeat titration two more times and complete table 2.

**Table 2**

|  |  |  |  |
| --- | --- | --- | --- |
|  | I | II | III |
| Final burette reading |  |  |  |
| Initial burette reading |  |  |  |
| Volume of solution C used (cm3) |  |  |  |

(4 marks)

Calculate the;

a) Average volume of sodium hydroxide used. (1 mark)

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b) The number of moles of;

I. Sodium hydroxide used. (1 mark)

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II) Hydrochloric acid in 25cm3 of solution V. (1 mark)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

III) Hydrochloric acid in 250cm3 Solution V. (1 mark)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

IV) Hydrochloric acid in 20.0 cm3 of solution V. (1 mark)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

V) Hydrochloric acid that reacted with solid V. (1 mark)

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

c) Calculate the enthalpy of reaction between solid V and one mole of hydrochloric acid. (2 marks)

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1. You are provided with mixture N. You are required to:-

i) Carry out tests on mixture N

ii) Identify any gases produced if any

iii) Record your observations and inferences accordingly.

**Procedure**

1. Place a spatulaful of mixture N in a test tube.
2. Add 8cm3 of distilled water and shake well
3. Filter and retain the residue

a) Divide the filtrate into four parts.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

b) Add sodium hydroxide to the first portion drop wise while observing till in excess.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

c) Add ammonia solution to the second portion of the filtrate drop wise, until in excess.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

d) Add 1cm3 dilute hydrochloric acid to the third portion.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

e) Add a few drops of potassium iodide solution to the fourth portion.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

f) Remove the residue from the filter paper and place it in a test tube, add 5cm3 of dilute nitric (V) acid.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

1. You are provided with organic compound solid G. Carry out the following tests.

a) Place all of solid G in a boiling tube. Add about 20cm3 of distilled water and shake well. Divide the mixture into 3 separate test tubes.

|  |  |
| --- | --- |
| Observations | Inferences |
| (½mark) | (½mark) |

b) To the first portion of the mixture add a spatula full of sodium carbonate solid.

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

c) To the second portion of the mixture, add a few drops of universal indicator and test the pH

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |

d) To the third portion of the mixture, add 2cm3 of ethanol followed by 2 drops of concentrated sulphuric (VI) acid

|  |  |
| --- | --- |
| Observations | Inferences |
| (1 mark) | (1 mark) |